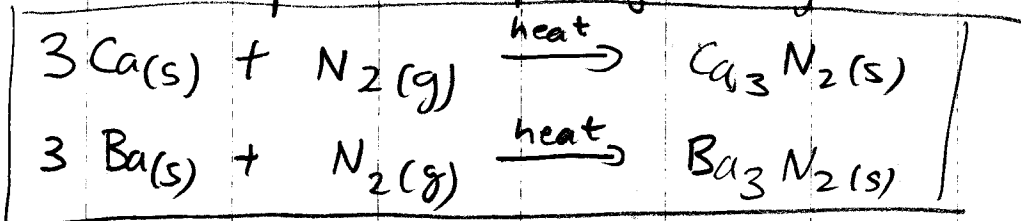


- 1.90 (a) Copper(II) cyanide  
 (b) Copper(I) carbonate  
 (c) Copper(II) phosphate  
 (d) Copper(II) sulfate  
 (e) Copper(I) sulfate

you should understand  
 where the (I) and (II)  
 oxidation states  
 come from!

1.92 Mg, Ca, and Ba are all Group 2.  
 So we'd predict completely analogous reactions:



1.110 (a) No carbon in compound (more or less).

(b) celestite: strontium sulfate

cinnabar: mercury(II) sulfide

corundum: aluminum oxide

cuprite: copper(I) oxide

fluorite: calcium fluoride

pyrolusite: manganese(IV) oxide

witherite: barium carbonate

zircon: zirconium(IV) silicate