
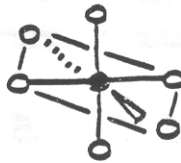
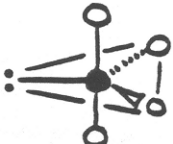


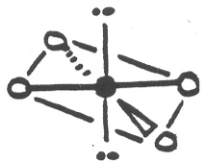



VSEPR Theory and Hypervalent Compounds

<u>S.N.</u>	<u>C.N.</u>	Molecular Shape (ideal bond angles)	<u>S.N.</u>	<u>C.N.</u>	Molecular Shape (ideal bond angles)
5	5	 <p>Trigonal bipyramidal 90°, 120°</p>	6	6	 <p>Octahedral 90°</p>
	4	 <p>See-saw 90°, 120°</p>		5	 <p>Square pyramidal 90°, 180°</p>
	3	 <p>T-shaped 90°</p>		4	 <p>Square planar 90°</p>
	2	 <p>Linear 180°</p>			

You are required to memorize the names of shapes SN = 2, 3, and 4 we have already discussed. However, you are not required to memorize the above names for SN = 5 and SN = 6 shapes. However, you should be able to explain the shapes and bond angles.