



Assume 100 mL of aqueous phase.

Say we can use 100 mL of CH_2Cl_2 for extraction

(1) 1 100-mL extraction

$$q = \frac{100 \text{ mL}}{100 \text{ mL} + 10(100 \text{ mL})} = 0.09 = \boxed{9\% \text{ remaining in aq phase}}$$

(2) 4 25-mL extractions

$$q^4 = \left(\frac{100 \text{ mL}}{100 \text{ mL} + 10(25 \text{ mL})} \right)^4 = 0.007 = \boxed{0.7\% \text{ remaining in aq phase}}$$

(3) 100 1-mL extractions

$$q^{100} = \left(\frac{100 \text{ mL}}{100 \text{ mL} + 10(1 \text{ mL})} \right)^{100} = 7 \times 10^{-5} = \boxed{0.007\% \text{ remaining in aq phase}}$$

* key is the chance to re-equilibrate with pure CH_2Cl_2 with each successive extraction